

Materials for electrochemical capacitors

Electrochemical capacitors, also called supercapacitors, store energy using either ion adsorption (electrochemical double layer capacitors) or fast surface redox reactions (pseudo-capacitors). They can complement or replace batteries in electrical energy storage and harvesting applications, when high power delivery or uptake is needed. A notable improvement in performance has been achieved through recent advances in understanding charge storage mechanisms and the development of advanced nanostructured materials. The discovery that ion desolvation occurs in pores smaller than the solvated ions has led to higher capacitance for electrochemical double layer capacitors using carbon electrodes with subnanometre pores, and opened the door to designing high-energy density devices using a variety of electrolytes. Combination of pseudo-capacitive nanomaterials, including oxides, nitrides and polymers, with the latest generation of nanostructured lithium electrodes has brought the energy density of electrochemical capacitors closer to that of batteries. The use of carbon nanotubes has further advanced micro-electrochemical capacitors, enabling flexible and adaptable devices to be made. Mathematical modelling and simulation will be the key to success in designing tomorrow's high-energy and high-power devices.

PATRICE SIMON^{1,2} AND YURY GOGOTSI³

¹Université Paul Sabatier, CIRIMAT, UMR-CNRS 5085, 31062 Toulouse Cedex 4, France

²Institut Universitaire de France, 103 Boulevard Saint Michel, 75005 Paris, France

³Department of Materials Science & Engineering, Drexel University, 3141 Chestnut Street, Philadelphia 19104, USA

e-mail: simon@chimie.ups-tlse.fr; gogotsi@drexel.edu

Climate change and the decreasing availability of fossil fuels require society to move towards sustainable and renewable resources. As a result, we are observing an increase in renewable energy production from sun and wind, as well as the development of electric vehicles or hybrid electric vehicles with low CO₂ emissions. Because the sun does not shine during the night, wind does not blow on demand and we all expect to drive our car with at least a few hours of autonomy, energy storage systems are starting to play a larger part in our lives. At the forefront of these are electrical energy storage systems, such as batteries and electrochemical capacitors (ECs)¹. However, we need to improve their performance substantially to meet the higher requirements of future systems, ranging from portable electronics to hybrid electric vehicles and large industrial equipment, by developing new materials and advancing our understanding of the electrochemical interfaces at the nanoscale. Figure 1 shows the plot of power against energy density, also called a Ragone plot², for the most important energy storage systems.

Lithium-ion batteries were introduced in 1990 by Sony, following pioneering work by Whittingham, Scrosati and Armand (see ref. 3 for a review). These batteries, although costly, are the best in terms of performance, with energy densities that can reach 180 watt hours

per kilogram. Although great efforts have gone into developing high-performance Li-ion and other advanced secondary batteries that use nanomaterials or organic redox couples⁴⁻⁶, ECs have attracted less attention until very recently. Because Li-ion batteries suffer from a somewhat slow power delivery or uptake, faster and higher-power energy storage systems are needed in a number of applications, and this role has been given to the ECs⁷. Also known as supercapacitors or ultracapacitors, ECs are power devices that can be fully charged or discharged in seconds; as a consequence, their energy density (about 5 Wh kg⁻¹) is lower than in batteries, but a much higher power delivery or uptake (10 kW kg⁻¹) can be achieved for shorter times (a few seconds)¹. They have had an important role in complementing or replacing batteries in the energy storage field, such as for uninterruptible power supplies (back-up supplies used to protect against power disruption) and load-levelling. A more recent example is the use of electrochemical double layer capacitors (EDLCs) in emergency doors (16 per plane) on an Airbus A380, thus proving that in terms of performance, safety and reliability ECs are definitely ready for large-scale implementation. A recent report by the US Department of Energy⁸ assigns equal importance to supercapacitors and batteries for future energy storage systems, and articles on supercapacitors appearing in business and popular magazines show increasing interest by the general public in this topic.

Several types of ECs can be distinguished, depending on the charge storage mechanism as well as the active materials used. EDLCs, the most common devices at present, use carbon-based active materials with high surface area (Fig. 2). A second group of ECs, known as pseudo-capacitors or redox supercapacitors, uses fast and reversible surface or near-surface reactions for charge storage. Transition metal oxides as well as electrically conducting polymers are examples of

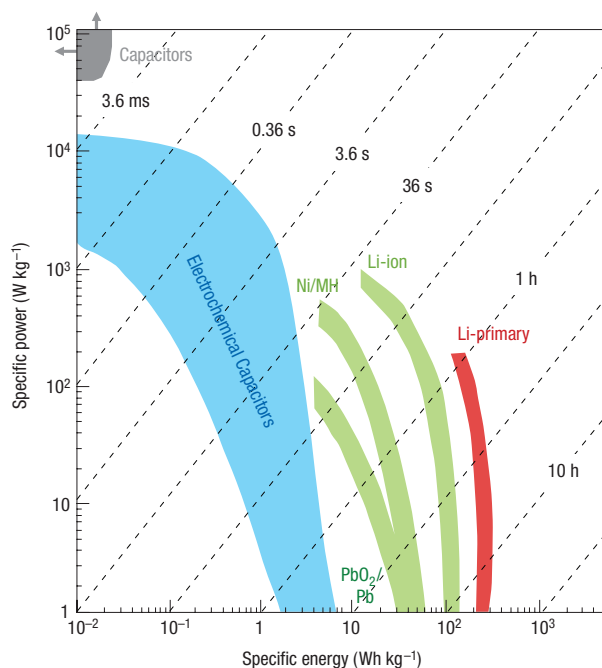


Figure 1 Specific power against specific energy, also called a Ragone plot, for various electrical energy storage devices. If a supercapacitor is used in an electric vehicle, the specific power shows how fast one can go, and the specific energy shows how far one can go on a single charge. Times shown are the time constants of the devices, obtained by dividing the energy density by the power.

pseudo-capacitive active materials. Hybrid capacitors, combining a capacitive or pseudo-capacitive electrode with a battery electrode, are the latest kind of EC, which benefit from both the capacitor and the battery properties.

Electrochemical capacitors currently fill the gap between batteries and conventional solid state and electrolytic capacitors (Fig. 1). They store hundreds or thousands of times more charge (tens to hundreds of farads per gram) than the latter, because of a much larger surface area ($1,000\text{--}2,000\text{ m}^2\text{ g}^{-1}$) available for charge storage in EDLC. However, they have a lower energy density than batteries, and this limits the optimal discharge time to less than a minute, whereas many applications clearly need more⁹. Since the early days of EC development in the late 1950s, there has not been a good strategy for increasing the energy density; only incremental performance improvements were achieved from the 1960s to 1990s. The impressive increase in performance that has been demonstrated in the past couple of years is due to the discovery of new electrode materials and improved understanding of ion behaviour in small pores, as well as the design of new hybrid systems combining faradic and capacitive electrodes. Here we give an overview of past and recent findings as well as an analysis of what the future holds for ECs.

ELECTROCHEMICAL DOUBLE-LAYER CAPACITORS

The first patent describing the concept of an electrochemical capacitor was filed in 1957 by Becker⁹, who used carbon with a high specific surface area (SSA) coated on a metallic current collector in a sulphuric acid solution. In 1971, NEC (Japan) developed aqueous-electrolyte capacitors under the energy company SOHIO's licence for power-saving units in electronics, and this application can be considered as the starting point for electrochemical capacitor use in commercial devices⁹. New applications in mobile electronics, transportation

(cars, trucks, trams, trains and buses), renewable energy production and aerospace systems¹⁰ bolstered further research.

MECHANISM OF DOUBLE-LAYER CAPACITANCE

EDLCs are electrochemical capacitors that store the charge electrostatically using reversible adsorption of ions of the electrolyte onto active materials that are electrochemically stable and have high accessible SSA. Charge separation occurs on polarization at the electrode–electrolyte interface, producing what Helmholtz described in 1853 as the double layer capacitance C :

$$C = \frac{\epsilon_r \epsilon_0 A}{d} \quad \text{or} \quad C/A = \frac{\epsilon_r \epsilon_0}{d} \quad (1)$$

where ϵ_r is the electrolyte dielectric constant, ϵ_0 is the dielectric constant of the vacuum, d is the effective thickness of the double layer (charge separation distance) and A is the electrode surface area.

This capacitance model was later refined by Gouy and Chapman, and Stern and Geary, who suggested the presence of a diffuse layer in the electrolyte due to the accumulation of ions close to the electrode surface. The double layer capacitance is between 5 and $20\text{ }\mu\text{F cm}^{-2}$ depending on the electrolyte used¹¹. Specific capacitance achieved with aqueous alkaline or acid solutions is generally higher than in organic electrolytes¹¹, but organic electrolytes are more widely used as they can sustain a higher operation voltage (up to 2.7 V in symmetric systems). Because the energy stored is proportional to voltage squared according to

$$E = \frac{1}{2} CV^2 \quad (2)$$

a three-fold increase in voltage, V , results in about an order of magnitude increase in energy, E , stored at the same capacitance.

As a result of the electrostatic charge storage, there is no faradic (redox) reaction at EDLC electrodes. A supercapacitor electrode must be considered as a blocking electrode from an electrochemical point of view. This major difference from batteries means that there is no limitation by the electrochemical kinetics through a polarization resistance. In addition, this surface storage mechanism allows very fast energy uptake and delivery, and better power performance. The absence of faradic reactions also eliminates the swelling in the active material that batteries show during charge/discharge cycles. EDLCs can sustain millions of cycles whereas batteries survive a few thousand at best. Finally, the solvent of the electrolyte is not involved in the charge storage mechanism, unlike in Li-ion batteries where it contributes to the solid–electrolyte interphase when graphite anodes or high-potential cathodes are used. This does not limit the choice of solvents, and electrolytes with high power performances at low temperatures (down to $-40\text{ }^\circ\text{C}$) can be designed for EDLCs. However, as a consequence of the electrostatic surface charging mechanism, these devices suffer from a limited energy density. This explains why today's EDLC research is largely focused on increasing their energy performance and widening the temperature limits into the range where batteries cannot operate⁹.

HIGH SURFACE AREA ACTIVE MATERIALS

The key to reaching high capacitance by charging the double layer is in using high SSA blocking and electronically conducting electrodes. Graphitic carbon satisfies all the requirements for this application, including high conductivity, electrochemical stability and open porosity¹². Activated, templated and carbide-derived carbons¹³, carbon fabrics, fibres, nanotubes¹⁴, onions¹⁵ and nanohorns¹⁶ have been tested for EDLC applications¹¹, and some of these carbons are shown in Fig. 2a–d. Activated carbons are the most widely used materials today, because of their high SSA and moderate cost.

Activated carbons are derived from carbon-rich organic precursors by carbonization (heat treatment) in inert atmosphere

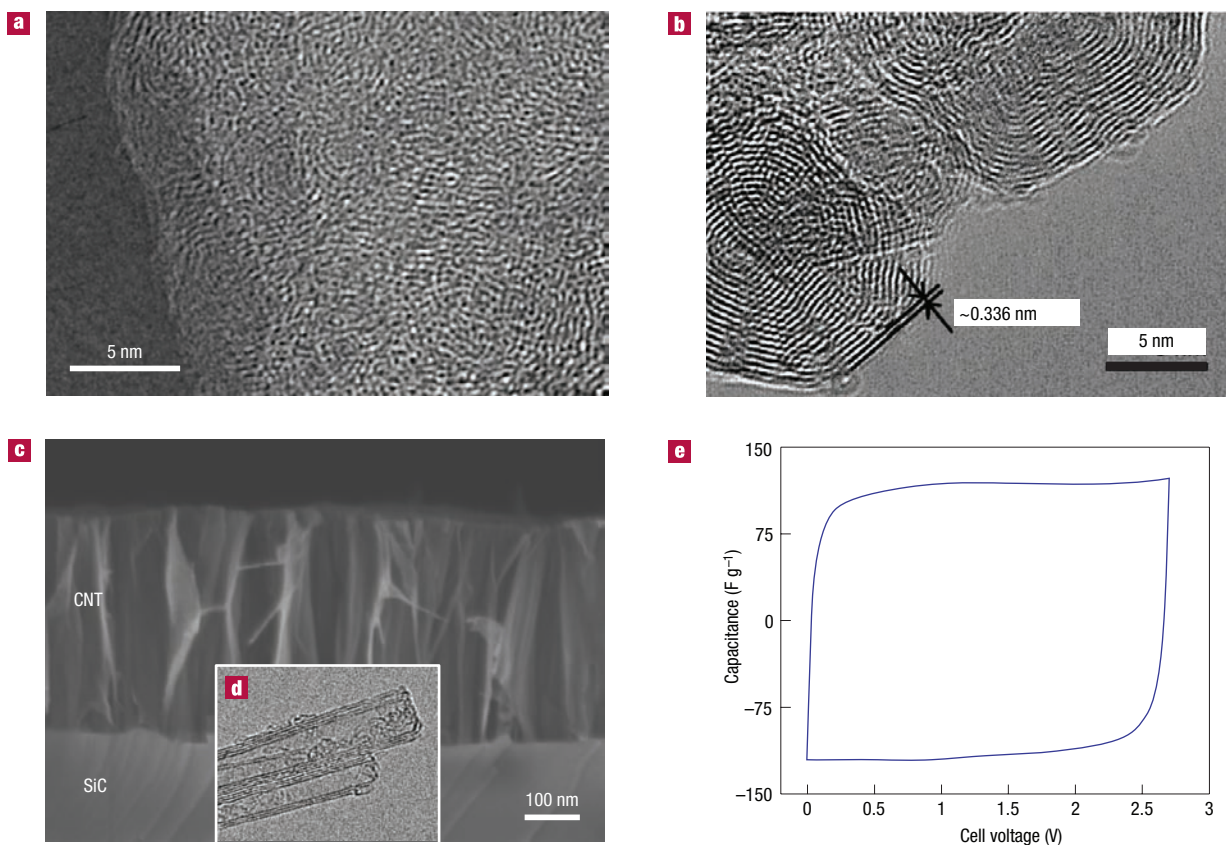


Figure 2 Carbon structures used as active materials for double layer capacitors. **a**, Typical transmission electronic microscopy (TEM) image of a disordered microporous carbon (SiC-derived carbon, 3 hours chlorination at 1,000 °C). **b**, TEM image of onion-like carbon. Reproduced with permission from ref. 80. © 2007 Elsevier. **c**, Scanning electron microscopy image of an array of carbon nanotubes (labelled CNT) on SiC produced by annealing for 6 h at 1,700 °C; inset, **d**, shows a TEM image of the same nanotubes⁷². **e**, Cyclic voltammetry of a two-electrode laboratory EDLC cell in 1.5 M tetraethylammonium tetrafluoroborate $\text{NET}_4^+\text{BF}_4^-$ in acetonitrile-based electrolyte, containing activated carbon powders coated on aluminium current collectors. Cyclic voltammetry was recorded at room temperature and potential scan rate of 20 mV s^{-1} .

with subsequent selective oxidation in CO_2 , water vapour or KOH to increase the SSA and pore volume. Natural materials, such as coconut shells, wood, pitch or coal, or synthetic materials, such as polymers, can be used as precursors. A porous network in the bulk of the carbon particles is produced after activation; micropores (<2 nm in size), mesopores (2–50 nm) and macropores (>50 nm) can be created in carbon grains. Accordingly, the porous structure of carbon is characterized by a broad distribution of pore size. Longer activation time or higher temperature leads to larger mean pore size. The double layer capacitance of activated carbon reaches 100–120 F g^{-1} in organic electrolytes; this value can exceed 150–300 F g^{-1} in aqueous electrolytes, but at a lower cell voltage because the electrolyte voltage window is limited by water decomposition. A typical cyclic voltammogram of a two-electrode EDLC laboratory cell is presented in Fig. 2e. Its rectangular shape is characteristic of a pure double layer capacitance mechanism for charge storage according to:

$$I = C \times \frac{dV}{dt} \quad (3)$$

where I is the current, (dV/dt) is the potential scan rate and C is the double layer capacitance. Assuming a constant value for C , for a given scan rate the current I is constant, as can be seen from Fig. 2e, where the cyclic voltammogram has a rectangular shape.

As previously mentioned, many carbons have been tested for EDLC applications and a recent paper¹¹ provides an overview of

what has been achieved. Untreated carbon nanotubes¹⁷ or nanofibres have a lower capacitance (around 50–80 F g^{-1}) than activated carbon in organic electrolytes. It can be increased up to 100 F g^{-1} or greater by grafting oxygen-rich groups, but these are often detrimental to cyclability. Activated carbon fabrics can reach the same capacitance as activated carbon powders, as they have similar SSA, but the high price limits their use to speciality applications. The carbons used in EDL capacitors are generally pre-treated to remove moisture and most of the surface functional groups present on the carbon surface to improve stability during cycling, both of which can be responsible for capacitance fading during capacitor ageing as demonstrated by Azais *et al.*¹⁸ using NMR and X-ray photoelectron spectroscopy techniques. Pandolfo *et al.*¹¹, in their review article, concluded that the presence of oxygenated groups also contributes to capacitor instability, resulting in an increased series resistance and deterioration of capacitance. Figure 3 presents a schematic of a commercial EDLC, showing the positive and the negative electrodes as well as the separator in rolled design (Fig. 3a,b) and flat design (button cell in Fig. 3c).

CAPACITANCE AND PORE SIZE

Initial research on activated carbon was directed towards increasing the pore volume by developing high SSA and refining the activation process. However, the capacitance increase was limited even for the most porous samples. From a series of activated carbons with different pore sizes in various electrolytes, it was shown that there was no linear relationship between the SSA and the capacitance^{19–21}. Some



Figure 3 Electrochemical capacitors. **a**, Schematic of a commercial spirally wound double layer capacitor. **b**, Assembled device weighing 500 g and rated for 2,600 F. (Photo courtesy of Batscap, Groupe Bolloré, France.) **c**, A small button cell, which is just 1.6 mm in height and stores 5 F. (Photo courtesy of Y-Carbon, US.) Both devices operate at 2.7 V.

studies suggested that pores smaller than 0.5 nm were not accessible to hydrated ions^{20,22} and that even pores under 1 nm might be too small, especially in the case of organic electrolytes, where the size of the solvated ions is larger than 1 nm (ref. 23). These results were consistent with previous work showing that ions carry a dynamic sheath of solvent molecules, the solvation shell²⁴, and that some hundreds of kilojoules per mole are required to remove it²⁵ in the case of water molecules. A pore size distribution in the range 2–5 nm, which is larger than the size of two solvated ions, was then identified as a way to improve the energy density and the power capability. Despite all efforts, only a moderate improvement has been made. Gravimetric capacitance in the range of 100–120 F g⁻¹ in organic and 150–200 F g⁻¹ in aqueous electrolytes has been achieved^{26,27} and ascribed to improved ionic mass transport inside mesopores. It was assumed that a well balanced micro- or mesoporosity (according to IUPAC classification, micropores are smaller than 2 nm, whereas mesopores are 2–50 nm in diameter) was needed to maximize capacitance²⁸.

Although fine-tuned mesoporous carbons failed to achieve high capacitance performance, several studies reported an important capacitive contribution from micropores. From experiments using activated carbon cloth, Salitra *et al.*²⁹ suggested that a partial desolvation of ions could occur, allowing access to small pores (<2 nm). High capacitance was observed for a mesoporous carbon containing large numbers of small micropores^{30–32}, suggesting that partial ion desolvation could lead to an improved capacitance. High capacitances (120 F g⁻¹ and 80 F cm⁻³) were found in organic electrolytes for microporous carbons (<1.5 nm)^{33,34}, contradicting the solvated ion adsorption theory. Using microporous activated coal-based carbon materials, Raymundo-Pinero *et al.*³⁵ observed the same effect and found a maximum capacitance for pore size at 0.7 and

0.8 nm for aqueous and organic electrolytes, respectively. However, the most convincing evidence of capacitance increase in pores smaller than the solvated ion size was provided by experiments using carbide-derived carbons (CDCs)^{36–38} as the active material. These are porous carbons obtained by extraction of metals from carbides (TiC, SiC and other) by etching in halogens at elevated temperatures³⁹:



In this reaction, Ti is leached out from TiC, and carbon atoms self-organize into an amorphous or disordered, mainly *sp*²-bonded⁴⁰, structure with a pore size that can be fine-tuned by controlling the chlorination temperature and other process parameters. Accordingly, a narrow uni-modal pore size distribution can be achieved in the range 0.6–1.1 nm, and the mean pore size can be controlled with sub-ångström accuracy⁴¹. These materials were used to understand the charge storage in micropores using 1 M solution of NEt₄BF₄ in acetonitrile-based electrolyte⁴². The normalized capacitance (μF cm⁻²) decreased with decreasing pore size until a critical value close to 1 nm was reached (Fig. 4), and then sharply increased when the pore size approached the ion size. As the CDC samples were exclusively microporous, the capacitance increase for subnanometre pores clearly shows the role of micropores. Moreover, the gravimetric and volumetric capacitances achieved by CDC were, respectively, 50% and 80% higher than for conventional activated carbon^{19–21}. The capacitance change with the current density was also found to be stable, demonstrating the high power capabilities these materials can achieve⁴². As the solvated ion sizes in this electrolyte were 1.3 and 1.16 nm for the cation and anion¹⁶, respectively, it was proposed that partial or complete removal of their solvation shell was allowing the ions to access the micropores. As a result, the change of capacitance was a linear function of 1/*b* (where *b* is the pore radius), confirming that the distance between the ion and the carbon surface, *d*, was shorter for the smaller pores. This dependence published by Chmiola *et al.*⁴² has since been confirmed by other studies, and analysis of literature data is provided in refs 43 and 44.

CHARGE-STORAGE MECHANISM IN SUBNANOMETRE PORES

From a fundamental point of view, there is a clear lack of understanding of the double layer charging in the confined space of micropores, where there is no room for the formation of the Helmholtz layer and diffuse layer expected at a solid–electrolyte interface. To address this issue, a three-electrode cell configuration, which discriminates between anion and cation adsorption, was used⁴⁵. The double layer capacitance in 1.5 M NEt₄BF₄-acetonitrile electrolyte caused by the anion and cation at the positive and negative electrodes, respectively, had maxima at different pore sizes⁴⁵. The peak in capacitance shifted to smaller pores for the smaller ion (anion). This behaviour cannot be explained by purely electrostatic reasons, because all pores in this study were the same size as or smaller than a single ion with a single associated solvent molecule. It thus confirmed that ions must be at least partially stripped of solvent molecules in order to occupy the carbon pores. These results point to a charge storage mechanism whereby partial or complete removal of the solvation shell and increased confinement of ions lead to increased capacitance.

A theoretical analysis published by Huang *et al.*⁴³ proposed splitting the capacitive behaviour in two different parts depending on the pore size. For mesoporous carbons (pores larger than 2 nm), the traditional model describing the charge of the double layer was used⁴³:

$$C/A = \frac{\epsilon_r \epsilon_0}{b \ln\left(\frac{b}{b-d}\right)} \quad (5)$$

where *b* is the pore radius and *d* is the distance of approach of the ion to the carbon surface. Data from Fig. 4 in the mesoporous range

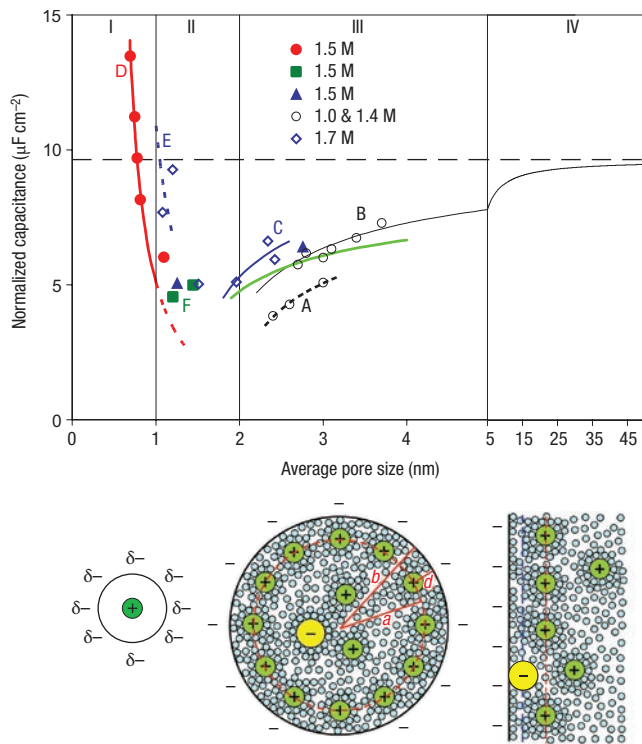


Figure 4 Specific capacitance normalized by SSA as a function of pore size for different carbon samples. All samples were tested in the same electrolyte ($\text{NET}_4^+, \text{BF}_4^-$ in acetonitrile; concentrations are shown in the key). Symbols show experimental data for CDCs, templated mesoporous carbons and activated carbons, and lines show model fits⁴³. A huge normalized capacitance increase is observed for microporous carbons with the smallest pore size in zone I, which would not be expected in the traditional view. The partial or complete loss of the solvation shell explains this anomalous behaviour⁴². As schematics show, zones I and II can be modelled as an electric wire-in-cylinder capacitor, an electric double-cylinder capacitor should be considered for zone III, and the commonly used planar electric double layer capacitor can be considered for larger pores, when the curvature/size effect becomes negligible (zone IV). A mathematical fit in the mesoporous range (zone III) is obtained using equation (5). Equation (6) was used to model the capacitive behaviour in zone I, where confined micropores force ions to desolvate partially or completely⁴⁴. A, B: templated mesoporous carbons; C: activated mesoporous carbon; D, F: microporous CDC; E: microporous activated carbon. Reproduced with permission from ref. 44. © 2008 Wiley.

(zone III) were fitted with equation (5). For micropores (<1 nm), it was assumed that ions enter a cylindrical pore and line up, thus forming the ‘electric wire in cylinder’ model of a capacitor. Capacitance was calculated from⁴³

$$C/A = \frac{\epsilon_r \epsilon_0}{b \ln \left(\frac{b}{a_0} \right)} \quad (6)$$

where a_0 is the effective size of the ion (desolvated). This model perfectly matches with the normalized capacitance change versus pore size (zone I in Fig. 4). Calculations using density functional theory gave consistent values for the size, a_0 , for unsolvated NET_4^+ and BF_4^- ions⁴³.

This work suggests that removal of the solvation shell is required for ions to enter the micropores. Moreover, the ionic radius a_0 found by using equation (6) was close to the bare ion size, suggesting that ions could be fully desolvated. A study carried out with CDCs in a

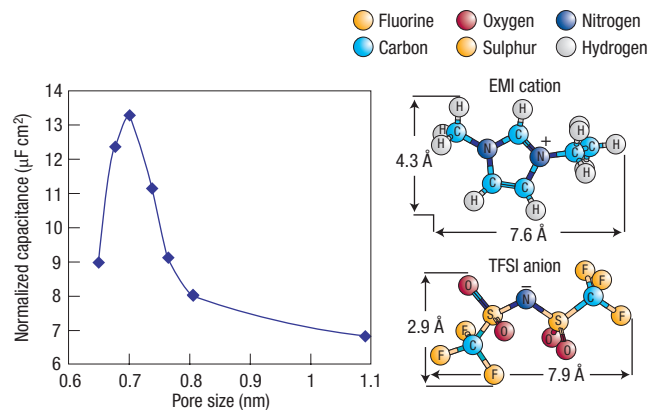


Figure 5 Normalized capacitance change as a function of the pore size of carbon-derived carbide samples. Samples were prepared at different temperatures in ethyl-methylimidazolium/trifluoro-methane-sulphonylimide (EMI,TFSI) ionic liquid at 60 °C. Inset shows the structure and size of the EMI and TFSI ions. The maximum capacitance is obtained when the pore size is in the same range as the maximum ion dimension. Reproduced with permission from ref. 46. © 2008 ACS.

solvent-free electrolyte ($[\text{EMI}^+, \text{TFSI}^-]$ ionic liquid at 60 °C), in which both ions have a maximum size of about 0.7 nm, showed the maximum capacitance for samples with the 0.7-nm pore size⁴⁶, demonstrating that a single ion per pore produces the maximum capacitance (Fig. 5). This suggests that ions cannot be adsorbed on both pore surfaces, in contrast with traditional supercapacitor models.

MATERIALS BY DESIGN

The recent findings of the micropore contribution to the capacitive storage highlight the lack of fundamental understanding of the electrochemical interfaces at the nanoscale and the behaviour of ions confined in nanopores. In particular, the results presented above rule out the generally accepted description of the double layer with solvated ions adsorbed on both sides of the pore walls, consistent with the absence of a diffuse layer in subnanometre pores. Although recent studies^{45,46} provide some guidance for developing materials with improved capacitance, such as elimination of macro- and mesopores and matching the pore size with the ion size, further material optimization by Edisonian or combinatorial electrochemistry methods may take a very long time. The effects of many parameters, such as carbon bonding (sp versus sp^2 or sp^3), pore shape, defects or adatoms, are difficult to determine experimentally. Clearly, computational tools and atomistic simulation will be needed to help us to understand the charge storage mechanism in subnanometre pores and to propose strategies to design the next generation of high-capacitance materials and material–electrolyte systems⁴⁷. Recasting the theory of double layers in electrochemistry to take into account solvation and desolvation effects could lead to a better understanding of charge storage as well as ion transport in ECs and even open up new opportunities in areas such as biological ion channels and water desalination.

REDOX-BASED ELECTROCHEMICAL CAPACITORS

MECHANISM OF PSEUDO-CAPACITIVE CHARGE STORAGE

Some ECs use fast, reversible redox reactions at the surface of active materials, thus defining what is called the pseudo-capacitive behaviour. Metal oxides such as RuO_2 , Fe_3O_4 or MnO_2 (refs 48, 49), as well as electronically conducting polymers⁵⁰, have been extensively studied in the past decades. The specific pseudo-capacitance exceeds

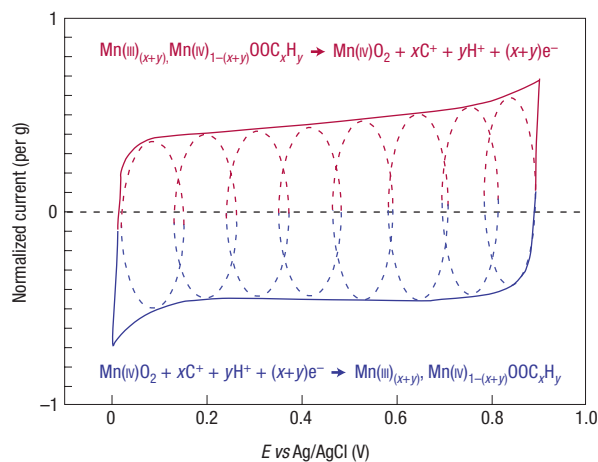
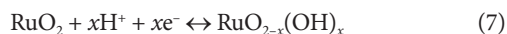


Figure 6 Cyclic voltammetry. This schematic of cyclic voltammetry for a MnO_2 -electrode cell in mild aqueous electrolyte (0.1 M K_2SO_4) shows the successive multiple surface redox reactions leading to the pseudo-capacitive charge storage mechanism. The red (upper) part is related to the oxidation from Mn(III) to Mn(IV) and the blue (lower) part refers to the reduction from Mn(IV) to Mn(III) .

that of carbon materials using double layer charge storage, justifying interest in these systems. But because redox reactions are used, pseudo-capacitors, like batteries, often suffer from a lack of stability during cycling.

Ruthenium oxide, RuO_2 , is widely studied because it is conductive and has three distinct oxidation states accessible within 1.2 V. The pseudo-capacitive behaviour of RuO_2 in acidic solutions has been the focus of research in the past 30 years¹. It can be described as a fast, reversible electron transfer together with an electro-adsorption of protons on the surface of RuO_2 particles, according to equation (7), where Ru oxidation states can change from (II) up to (IV):



where $0 \leq x \leq 2$. The continuous change of x during proton insertion or de-insertion occurs over a window of about 1.2 V and leads to a capacitive behaviour with ion adsorption following a Frumkin-type isotherm¹. Specific capacitance of more than 600 F g^{-1} has been reported⁵¹, but Ru-based aqueous electrochemical capacitors are expensive, and the 1-V voltage window limits their applications to small electronic devices. Organic electrolytes with proton surrogates (for example Li^+) must be used to go past 1 V. Less expensive oxides of iron, vanadium, nickel and cobalt have been tested in aqueous electrolytes, but none has been investigated as much as manganese oxide⁵². The charge storage mechanism is based on surface adsorption of electrolyte cations C^+ (K^+ , Na^+ ...) as well as proton incorporation according to the reaction:

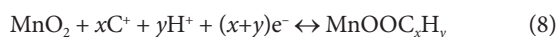


Figure 6 shows a cyclic voltammogram of a single MnO_2 electrode in mild aqueous electrolyte; the fast, reversible successive surface redox reactions define the behaviour of the voltammogram, whose shape is close to that of the EDLC. MnO_2 micro-powders or micrometre-thick films show a specific capacitance of about 150 F g^{-1} in neutral aqueous electrolytes within a voltage window of $<1 \text{ V}$. Accordingly, there is limited interest in MnO_2 electrodes for symmetric devices, because there are no oxidation states available at less than 0 V. However, it is suitable for a pseudo-capacitive positive electrode in hybrid systems,

which we will describe below. Other transition metal oxides with various oxidation degrees, such as molybdenum oxides, should also be explored as active materials for pseudo-capacitors.

Many kinds of conducting polymers (polyaniline, polypyrrole, polythiophene and their derivatives) have been tested in EC applications as pseudo-capacitive materials^{50,53,54} and have shown high gravimetric and volumetric pseudo-capacitance in various non-aqueous electrolytes at operating voltages of about 3 V. When used as bulk materials, conducting polymers suffer from a limited stability during cycling that reduces the initial performance⁹. Research efforts with conducting polymers for supercapacitor applications are nowadays directed towards hybrid systems.

NANOSTRUCTURING REDOX-ACTIVE MATERIALS TO INCREASE CAPACITANCE

Given that nanomaterials have helped to improve Li-ion batteries⁵⁵, it is not surprising that nanostructuring has also affected ECs. Because pseudo-capacitors store charge in the first few nanometres from the surface, decreasing the particle size increases active material usage. Thanks to a thin electrically conducting surface layer of oxide and oxynitride, the charging mechanism of nanocrystalline vanadium nitride (VN) includes a combination of an electric double layer and a faradic reaction (II/IV) at the surface of the nanoparticles, leading to specific capacitance up to $1,200 \text{ F g}^{-1}$ at a scan rate of 2 mV s^{-1} (ref. 56). A similar approach can be applied to other nano-sized transition metal nitrides or oxides. In another example, the cycling stability and the specific capacitance of RuO_2 nanoparticles were increased by depositing a thin conducting polymer coating that enhanced proton exchange at the surface⁵⁷. The design of specific surface functionalization to improve interfacial exchange could be suggested as a generic approach to other pseudo-redox materials.

MnO_2 and RuO_2 films have been synthesized at the nanometre scale. Thin MnO_2 deposits of tens to hundreds of nanometres have been produced on various substrates such as metal collectors, carbon nanotubes or activated carbons. Specific capacitances as high as $1,300 \text{ F g}^{-1}$ have been reported⁵⁸, as reaction kinetics were no longer limited by the electrical conductivity of MnO_2 . In the same way, Sugimoto's group have prepared hydrated RuO_2 nano-sheets with capacitance exceeding $1,300 \text{ F g}^{-1}$ (ref. 59). The RuO_2 specific capacitance also increased sharply when the film thickness was decreased. The deposition of RuO_2 thin film onto carbon supports^{60,61} both increased the capacitance and decreased the RuO_2 consumption. Thin film synthesis or high SSA capacitive material decoration with nano-sized pseudo-capacitive active material, like the examples presented in Fig. 7a and b, offers an opportunity to increase energy density and compete with carbon-based EDLCs. Particular attention must be paid to further processing of nano-sized powders into active films because they tend to re-agglomerate into large-size grains. An alternative way to produce porous films from powders is by growing nanotubes, as has been shown for V_2O_5 (ref. 62), or nanorods. These allow easy access to the active material, but can only be produced in thin films so far, and the manufacturing cost will probably limit the use of these sophisticated nanostructures to small electronic devices.

HYBRID SYSTEMS TO ACHIEVE HIGH ENERGY DENSITY

Hybrid systems offer an attractive alternative to conventional pseudo-capacitors or EDLCs by combining a battery-like electrode (energy source) with a capacitor-like electrode (power source) in the same cell. An appropriate electrode combination can even increase the cell voltage, further contributing to improvement in energy and power densities. Currently, two different approaches to hybrid systems have emerged: (i) pseudo-capacitive metal oxides with a capacitive carbon electrode, and (ii) lithium-insertion electrodes with a capacitive carbon electrode.

Numerous combinations of positive and negative electrodes have been tested in the past in aqueous or inorganic electrolytes. In most

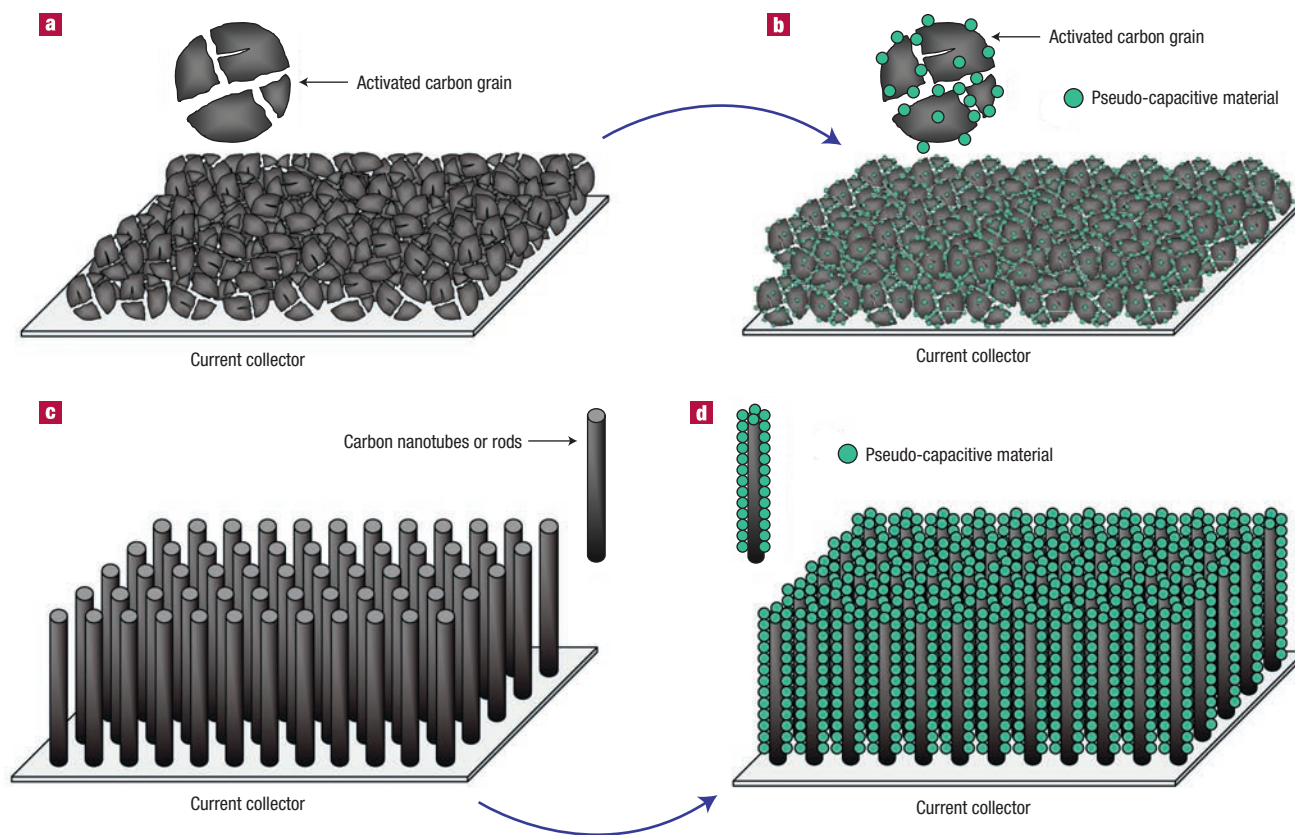


Figure 7 Possible strategies to improve both energy and power densities for electrochemical capacitors. **a, b**, Decorating activated carbon grains **(a)** with pseudo-capacitive materials **(b)**. **c, d**, Achieving conformal deposit of pseudo-capacitive materials **(d)** onto highly ordered high-surface-area carbon nanotubes **(c)**.

cases, the faradic electrode led to an increase in the energy density at the cost of cyclability (for balanced positive and negative electrode capacities). This is certainly the main drawback of hybrid devices, as compared with EDLCs, and it is important to avoid transforming a good supercapacitor into a mediocre battery⁶³.

MnO₂ is one of the most studied materials as a low-cost alternative to RuO₂. Its pseudo-capacitance arises from the III/IV oxidation state change at the surface of MnO₂ particles⁵⁸. The association of a negative EDLC-type electrode with a positive MnO₂ electrode leads to a 2-V cell in aqueous electrolytes thanks to the apparent water decomposition overvoltage on MnO₂ and high-surface-area carbon. The low-cost carbon–MnO₂ hybrid system combines high capacitance in neutral aqueous electrolytes with high cell voltages, making it a green alternative to EDLCs using acetonitrile-based solvents and fluorinated salts. Moreover, the use of MnO₂ nano-powders and nanostructures offers the potential for further improvement in capacitance⁶⁴. Another challenge for this system is to use organic electrolytes to reach higher cell voltage, thus improving the energy density.

A combination of a carbon electrode with a PbO₂ battery-like electrode using H₂SO₄ solution can work at 2.1 V (ref. 65), offering a low-cost EC device for cost-sensitive applications, in which weight of the device is of minor concern.

The hybrid concept originated from the Li-ion batteries field. In 1999, Amatucci's group combined a nanostructured lithium titanate anode Li₄Ti₅O₁₂ with an activated carbon positive electrode, designing a 2.8-V system that for the first time exceeded 10 Wh kg⁻¹ (ref. 66). The titanate electrode ensured high power capacity and no solid-electrolyte interphase formation, as well as long-life cyclability thanks to low volume change during cycling. Following this pioneering

work, many studies have been conducted on various combinations of a lithium-insertion electrode with a capacitive carbon electrode. The Li-ion capacitor developed by Fuji Heavy Industry is an example of this concept, using a pre-lithiated high SSA carbon anode together with an activated carbon cathode^{63,67}. It achieved an energy density of more than 15 Wh kg⁻¹ at 3.8 V. Capacity retention was increased by unbalancing the electrode capacities, allowing a low depth of charge/discharge at the anode. Systems with an activated carbon anode and anion intercalation cathode are also under development. The advent of nanomaterials⁵⁵ as well as fast advances in the area of Li-ion batteries should lead to the design of high-performance ECs. Combining newly developed high-rate conversion reaction anodes or Li-alloying anodes with a positive supercapacitor electrode could fill the gap between Li-ion batteries and EDLCs. These systems could be of particular interest in applications where high power and medium cycle life are needed.

CURRENT COLLECTORS

Because ECs are power devices, their internal resistance must be kept low. Particular attention must be paid to the contact impedance between the active film and the current collector. ECs designed for organic electrolytes use treated aluminium foil or grid current collectors. Surface treatments have already been shown to decrease ohmic drops at this interface⁶⁸, and coatings on aluminium that improve electrochemical stability at high potentials and interface conductivity are of great interest.

The design of nanostructured current collectors with an increased contact area is another way to control the interface between current

collector and active material. For example, carbon can be produced in a variety of morphologies¹², including porous films and nanotube brushes that can be grown on various current collectors⁶⁹ and that can serve as substrates for further conformal deposition (Fig. 7c and d) of active material. These nano-architected electrodes could outperform the existing systems by confining a highly pseudo-capacitive material to a thin film with a high SSA, as has been done for Li-ion batteries⁷⁰ where, by growing Cu nano-pillars on a planar Cu foil, a six-fold improvement in the energy density over planar electrodes has been achieved⁷⁰. Long's group⁶⁴ successfully applied a similar approach to supercapacitors by coating a porous carbon nano-foam with a 20-nm pseudo-capacitive layer of MnO₂. As a result, the area-normalized capacitance doubled to reach 1.5 F cm⁻², together with an outstanding volumetric capacitance of 90 F cm⁻³. Electrophoretic deposition from stable colloidal suspensions of RuO₂ (ref. 71) or other active material can be used for filling the inter-tube space to design high-energy-density devices which are just a few micrometres thick. The nano-architected electrodes also find applications in micro-systems where micro-ECs can complement micro-batteries for energy harvesting or energy generation. In this specific field, it is often advantageous to grow self-supported, binderless nano-electrodes directly on semiconductor wafers, such as Si or SiC (ref. 72; Fig. 2c).

An attractive material for current collectors is carbon in the form of a highly conductive nanotube or graphene paper. It does not corrode in aqueous electrolytes and is very flexible. The use of nanotube paper for manufacturing flexible supercapacitors is expected to grow as the cost of small-diameter nanotubes required for making paper decreases. The same thin sheet of nanotubes¹⁴ could potentially act as an active material and current collector at the same time. Thin-film, printable and wearable ECs could find numerous applications.

FROM ORGANIC TO IONIC LIQUID ELECTROLYTES

EC cell voltage is limited by the electrolyte decomposition at high potentials. Accordingly, the larger the electrolyte stability voltage window, the higher the supercapacitor cell voltage. Moving from aqueous to organic electrolytes increased the cell voltage from 0.9 V to 2.5–2.7 V for EDLCs. Because the energy density is proportional to the voltage squared (equation (2)), numerous research efforts have been directed at the design of highly conducting, stable electrolytes with a wider voltage window. Today, the state of the art is the use of organic electrolyte solutions in acetonitrile or propylene carbonate, the latter becoming more popular because of the low flash point and lower toxicity compared with acetonitrile.

Ionic liquids are room-temperature liquid solvent-free electrolytes; their voltage window stability is thus only driven by the electrochemical stability of the ions. A careful choice of both the anion and the cation allows the design of high-voltage supercapacitors, and 3-V, 1,000-F commercial devices are already available⁷³. However, the ionic conductivity of these liquids at room temperature is just a few milliSiemens per centimetre, so they are mainly used at higher temperatures. For example, CDC with an EMI/TFSI ionic liquid electrolyte has been shown⁴⁶ to have capacitance of 160 F g⁻¹ and ~90 F cm⁻³ at 60 °C. In this area, hybrid activated carbon/conducting polymer devices also show an improved performance with cell voltages higher than 3 V (refs 74–76).

For applications in the temperature range –30 °C to +60 °C, where batteries and supercapacitors are mainly used, ionic liquids still fail to satisfy the requirements because of their low ionic conductivity. However, the choice of a huge variety of combinations of anions and cations offers the potential for designing an ionic liquid electrolyte with an ionic conductivity of 40 mS cm⁻¹ and a voltage window of >4 V at room temperature⁷⁷. A challenge is, for instance, to find an alternative to the imidazolium cation that, despite high conductivity,

undergoes a reduction reaction at potential <1.5 V versus Li⁺/Li. Replacing the heavy bis(trifluoromethanesulphonyl)imide (TFSI) anion by a lighter (fluoromethanesulphonyl)imide (FSI) and preparing ionic liquid eutectic mixtures would improve both the cell voltage (because a protecting layer of AlF₃ can be formed on the Al surface, shifting the de-passivation potential of Al above 4 V) and the ionic conductivity⁷⁷. However, FSI shows poor cyclability at elevated temperatures. Supported by the efforts of the Li-ion community to design safer systems using ionic liquids, the research on ionic liquids for ECs is expected to have an important role in the improvement of capacitor performance in the coming years.

APPLICATIONS OF ELECTROCHEMICAL CAPACITORS

ECs are electrochemical energy sources with high power delivery and uptake, with an exceptional cycle life. They are used when high power demands are needed, such as for power buffer and power saving units, but are also of great interest for energy recovery. Recent articles from Miller *et al.*^{7,10} present an overview of the opportunities for ECs in a variety of applications, complementing an earlier review by Kötz *et al.*⁹. Small devices (a few farads) are widely used for power buffer applications or for memory back-up in toys, cameras, video recorders, mobile phones and so forth. Cordless tools such as screwdrivers and electric cutters using EDLCs are already available on the market. Such systems, using devices of a few tens of farads, can be fully charged or discharged in less than 2 minutes, which is particularly suited to these applications, with the cycle life of EDLC exceeding that of the tool. As mentioned before, the Airbus A380 jumbo jets use banks of EDLCs for emergency door opening. The modules consist of an in series/parallel assembly of 100-F, 2.7-V cells that are directly integrated into the doors to limit the use of heavy copper cables. This application is obviously a niche market, but it is a demonstration that the EDLC technology is mature in terms of performance, reliability and safety.

The main market targeted by EDLC manufacturers for the next years is the transportation market, including hybrid electric vehicles, as well as metro trains and tramways. There continues to be debate about the advantage of using high power Li-ion batteries instead of ECs (or vice versa) for these applications. Most of these discussions have been initiated by Li-ion battery manufacturers who would like their products to cover the whole range of applications. However, ECs and Li-ion batteries should not necessarily be seen as competitors, because their charge storage mechanisms and thus their characteristics are different. The availability of the stored charge will always be faster for a supercapacitor (surface storage) than for a Li-ion battery (bulk storage), with a larger stored energy for the latter. Both devices must be used in their respective time-constant domains (see Fig. 1). Using a Li-ion battery for repeated high power delivery/uptake applications for a short duration (10 s or less) will quickly degrade the cycle life of the system¹⁰. The only way to avoid this is to oversize the battery, increasing the cost and volume. In the same way, using ECs for power delivery longer than 10 s requires oversizing. However, some applications use ECs as the main power and energy source, benefiting from the fast charge/discharge capability of these systems as well as their outstanding cycle life. Several train manufacturers have clearly identified the tramway/metro market segment as extremely relevant for EC use, to power trains over short distances in big cities, where electric cables are clearly undesirable for aesthetic and other reasons, but also to recover the braking energy of another train on the same line, thanks to the ECs' symmetric high power delivery/uptake characteristics.

For automotive applications, manufacturers are already proposing solutions for electrical power steering, where ECs are used for load-leveling in stop-and-go traffic⁷⁸. The general trend is to increase the hybridization degree of the engines in hybrid electric vehicles, to allow fast acceleration (boost) and braking energy recovery. The

on-board energy storage systems will be in higher demand, and a combination of batteries and EDLCs will increase the battery cycle life, explaining why EDLCs are viewed as a partner to Li-ion batteries for this market⁷⁸. Currently, high price limits the use of both Li-ion batteries and EDLC in large-scale applications (for example for load levelling). But the surprisingly high cost of materials used for EDLC is due to a limited number of suppliers rather than intrinsically high cost of porous carbon. Decreasing the price of carbon materials for ECs, including CDC and AC, would remove the main obstacle to their wider use⁷⁹.

SUMMARY AND OUTLOOK

The most recent advances in supercapacitor materials include nanoporous carbons with the pore size tuned to fit the size of ions of the electrolyte with ångström accuracy, carbon nanotubes for flexible and printable devices with a short response time, and transition metal oxide and nitride nanoparticles for pseudo-capacitors with a high energy density. An improved understanding of charge storage and ion desolvation in subnanometre pores has helped to overcome a barrier that has been hampering progress in the field for decades. It has also shown how important it is to match the active materials with specific electrolytes and to use a cathode and anode with different pore sizes that match the anion or cation size. Nano-architecture of electrodes has led to further improvements in power delivery. The very large number of possible active materials and electrolytes means that better theoretical guidance is needed for the design of future ECs.

Future generations of ECs are expected to come close to current Li-ion batteries in energy density, maintaining their high power density. This may be achieved by using ionic liquids with a voltage window of more than 4 V, by discovering new materials that combine double-layer capacitance and pseudo-capacitance, and by developing hybrid devices. ECs will have a key role in energy storage and harvesting, decreasing the total energy consumption and minimizing the use of hydrocarbon fuels. Capacitive energy storage leads to a lower energy loss (higher cycle efficiency), than for batteries, compressed air, flywheel or other devices, helping to improve storage economy further. Flexible, printable and wearable ECs are likely to be integrated into smart clothing, sensors, wearable electronics and drug delivery systems. In some instances they will replace batteries, but in many cases they will either complement batteries, increasing their efficiency and lifetime, or serve as energy solutions where an extremely large number of cycles, long lifetime and fast power delivery are required.

doi:10.1038/nmat2297

References

- Conway, B. E. *Electrochemical Supercapacitors: Scientific Fundamentals and Technological Applications* (Kluwer, 1999).
- Service, R. F. New 'supercapacitor' promises to pack more electrical punch. *Science* **313**, 902–905 (2006).
- Tarascon, J.-M. & Armand, M. Issues and challenges facing rechargeable lithium batteries. *Nature* **414**, 359–367 (2001).
- Brodd, R. J. et al. Batteries, 1977 to 2002. *J. Electrochem. Soc.* **151**, K1–K11 (2004).
- Armand, M. & Tarascon, J.-M. Building better batteries. *Nature* **451**, 652–657 (2008).
- Armand, M. & Johansson, P. Novel weakly coordinating heterocyclic anions for use in lithium batteries. *J. Power Sources* **178**, 821–825 (2008).
- Miller, J. R. & Simon, P. Electrochemical capacitors for energy management. *Science* **321**, 651–652 (2008).
- US Department of Energy. *Basic Research Needs for Electrical Energy Storage* <www.sc.doe.gov/bes-reports/abstracts.html#EES2007> (2007).
- Kötz, R. & Carlen, M. Principles and applications of electrochemical capacitors. *Electrochim. Acta* **45**, 2483–2498 (2000).
- Miller, J. R. & Burke, A. F. Electrochemical capacitors: Challenges and opportunities for real-world applications. *Electrochem. Soc. Interf.* **17**, 53–57 (2008).
- Pandolfo, A. G. & Hollenkamp, A. F. Carbon properties and their role in supercapacitors. *J. Power Sources* **157**, 11–27 (2006).
- Gogotsi, Y. (ed.) *Carbon Nanomaterials* (CRC, 2006).
- Kyotani, T., Chmiola, J. & Gogotsi, Y. in *Carbon Materials for Electrochemical Energy Storage Systems* (eds Beguin, F. & Frackowiak, E.) Ch. 13 (CRC/Taylor and Francis, in the press).

- Futaba, D. N. et al. Shape-engineerable and highly densely packed single-walled carbon nanotubes and their application as super-capacitor electrodes. *Nature Mater.* **5**, 987–994 (2006).
- Portet, C., Chmiola, J., Gogotsi, Y., Park, S. & Lian, K. Electrochemical characterizations of carbon nanomaterials by the cavity microelectrode technique. *Electrochim. Acta*, **53**, 7675–7680 (2008).
- Yang, C.-M. et al. Nanowindow-regulated specific capacitance of supercapacitor electrodes of single-wall carbon nanohorns. *J. Am. Chem. Soc.* **129**, 20–21 (2007).
- Niu, C., Sichel, E. K., Hoch, R., Moy, D. & Tennent, H. High power electrochemical capacitors based on carbon nanotube electrodes. *Appl. Phys. Lett.* **70**, 1480 (1997).
- Azaïs, P. et al. Causes of supercapacitors ageing in organic electrolyte. *J. Power Sources* **171**, 1046–1053 (2007).
- Gamby, J., Taberna, P. L., Simon, P., Fauvarque, J. F. & Chesneau, M. Studies and characterization of various activated carbons used for carbon/carbon supercapacitors. *J. Power Sources* **101**, 109–116 (2001).
- Shi, H. Activated carbons and double layer capacitance. *Electrochim. Acta* **41**, 1633–1639 (1995).
- Qu, D. & Shi, H. Studies of activated carbons used in double-layer capacitors. *J. Power Sources* **74**, 99–107 (1998).
- Qu, D. Studies of the activated carbons used in double-layer supercapacitors. *J. Power Sources* **109**, 403–411 (2002).
- Kim, Y. J. et al. Correlation between the pore and solvated ion size on capacitance uptake of PVDC-based carbons. *Carbon* **42**, 1491 (2004).
- Izutsu, K. *Electrochemistry in Nonaqueous Solution* (Wiley, 2002).
- Marcus, Y. *Ion Solvation* (Wiley, 1985).
- Jurewicz, K. et al. Capacitance properties of ordered porous carbon materials prepared by a templating procedure. *J. Phys. Chem. Solids* **65**, 287 (2004).
- Fernández, J. A. et al. Performance of mesoporous carbons derived from poly(vinyl alcohol) in electrochemical capacitors. *J. Power Sources* **175**, 675 (2008).
- Fuertes, A. B., Lota, G., Centeno, T. A. & Frackowiak, E. Templated mesoporous carbons for supercapacitor application. *Electrochim. Acta* **50**, 2799 (2005).
- Salitra, G., Soffer, A., Eliad, L., Cohen, Y. & Aurbach, D. Carbon electrodes for double-layer capacitors. I. Relations between ion and pore dimensions. *J. Electrochem. Soc.* **147**, 2486–2493 (2000).
- Vix-Guterl, C. et al. Electrochemical energy storage in ordered porous carbon materials. *Carbon* **43**, 1293–1302 (2005).
- Eliad, L., Salitra, G., Soffer, A. & Aurbach, D. On the mechanism of selective electroadsorption of protons in the pores of carbon molecular sieves. *Langmuir* **21**, 3198–3202 (2005).
- Eliad, L. et al. Assessing optimal pore-to-ion size relations in the design of porous poly(vinylidene chloride) carbons for EDL capacitors. *Appl. Phys. A* **82**, 607–613 (2006).
- Arulepp, M. et al. The advanced carbide-derived carbon based supercapacitor. *J. Power Sources* **162**, 1460–1466 (2006).
- Arulepp, M. et al. Influence of the solvent properties on the characteristics of a double layer capacitor. *J. Power Sources* **133**, 320–328 (2004).
- Raymundo-Pinero, E., Kierzek, K., Machnikowski, J. & Beguin, F. Relationship between the nanoporous texture of activated carbons and their capacitance properties in different electrolytes. *Carbon* **44**, 2498–2507 (2006).
- Janes, A. & Lust, E. Electrochemical characteristics of nanoporous carbide-derived carbon materials in various nonaqueous electrolyte solutions. *J. Electrochem. Soc.* **153**, A113–A116 (2006).
- Shanina, B. D. et al. A study of nanoporous carbon obtained from ZC powders (Z = Si, Ti, and B). *Carbon* **41**, 3027–3036 (2003).
- Chmiola, J., Dash, R., Yushin, G. & Gogotsi, Y. Effect of pore size and surface area of carbide derived carbon on specific capacitance. *J. Power Sources* **158**, 765–772 (2006).
- Dash, R. et al. Titanium carbide derived nanoporous carbon for energy-related applications. *Carbon* **44**, 2489–2497 (2006).
- Urbanaitė, S. et al. EELS studies of carbide derived carbons. *Carbon* **45**, 2047–2053 (2007).
- Gogotsi, Y. et al. Nanoporous carbide-derived carbon with tunable pore size. *Nature Mater.* **2**, 591–594 (2003).
- Chmiola, J. et al. Anomalous increase in carbon capacitance at pore size below 1 nm. *Science* **313**, 1760–1763 (2006).
- Huang, J. S., Sumpter, B. G. & Meunier, V. Theoretical model for nanoporous carbon supercapacitors. *Angew. Chem. Int. Ed.* **47**, 520–524 (2008).
- Huang, J., Sumpter, B. G. & Meunier, V. A universal model for nanoporous carbon supercapacitors applicable to diverse pore regimes, carbons, and electrolytes. *Chem. Eur. J.* **14**, 6614–6626 (2008).
- Chmiola, J., Largeot, C., Taberna, P.-L., Simon, P. & Gogotsi, Y. Desolvation of ions in subnanometer pores, its effect on capacitance and double-layer theory. *Angew. Chem. Int. Ed.* **47**, 3392–3395 (2008).
- Largeot, C. et al. Relation between the ion size and pore size for an electric double-layer capacitor. *J. Am. Chem. Soc.* **130**, 2730–2731 (2008).
- Weigand, G., Davenport, J. W., Gogotsi, Y. & Roberto, J. in *Scientific Impacts and Opportunities for Computing* Ch. 5, 29–35 (DOE Office of Science, 2008).
- Wu, N.-L. Nanocrystalline oxide supercapacitors. *Mater. Chem. Phys.* **75**, 6–11 (2002).
- Brousse, T. et al. Crystalline MnO₂ as possible alternatives to amorphous compounds in electrochemical supercapacitors. *J. Electrochem. Soc.* **153**, A2171–A2180 (2006).
- Rudge, A., Raistrick, I., Gottesfeld, S. & Ferraris, J. P. Conducting polymers as active materials in electrochemical capacitors. *J. Power Sources* **47**, 89–107 (1994).
- Zheng, J. P. & Jow, T. R. High energy and high power density electrochemical capacitors. *J. Power Sources* **62**, 155–159 (1996).
- Lee, H. Y. & Goodenough, J. B. Supercapacitor behavior with KCl electrolyte. *J. Solid State Chem.* **144**, 220–223 (1999).
- Laforge, A., Simon, P. & Fauvarque, J.-F. Chemical synthesis and characterization of fluorinated polyphenylthiophenes: application to energy storage. *Synth. Met.* **123**, 311–319 (2001).
- Naoi, K., Suematsu, S. & Manago, A. Electrochemistry of poly(1,5-diaminoanthraquinone) and its application in electrochemical capacitor materials. *J. Electrochem. Soc.* **147**, 420–426 (2000).
- Arico, A. S., Bruce, P., Scrosati, B., Tarascon, J.-M. & Schalkwijk, W. V. Nanostructured materials for advanced energy conversion and storage devices. *Nature Mater.* **4**, 366–377 (2005).
- Choi, D., Blomgren, G. E. & Kumta, P. N. Fast and reversible surface redox reaction in nanocrystalline vanadium nitride supercapacitors. *Adv. Mater.* **18**, 1178–1182 (2006).

57. Machida, K., Furuuchi, K., Min, M. & Naoi, K. Mixed proton–electron conducting nanocomposite based on hydrous RuO₂ and polyaniline derivatives for supercapacitors. *Electrochemistry* **72**, 402–404 (2004).
58. Toupin, M., Brousse, T. & Belanger, D. Charge storage mechanism of MnO₂ electrode used in aqueous electrochemical capacitor. *Chem. Mater.* **16**, 3184–3190 (2004).
59. Sugimoto, W., Iwata, H., Yasunaga, Y., Murakami, Y. & Takasu, Y. Preparation of ruthenic acid nanosheets and utilization of its interlayer surface for electrochemical energy storage. *Angew. Chem. Int. Ed.* **42**, 4092–4096 (2003).
60. Miller, J. M., Dunn, B., Tran, T. D. & Pekala, R. W. Deposition of ruthenium nanoparticles on carbon aerogels for high energy density supercapacitor electrodes. *J. Electrochem. Soc.* **144**, L309–L311 (1997).
61. Min, M., Machida, K., Jang, J. H. & Naoi, K. Hydrous RuO₂/carbon black nanocomposites with 3D porous structure by novel incipient wetness method for supercapacitors. *J. Electrochem. Soc.* **153**, A334–A338 (2006).
62. Wang, Y., Takahashi, K., Lee, K. H. & Cao, G. Z. Nanostructured vanadium oxide electrodes for enhanced lithium-ion intercalation. *Adv. Funct. Mater.* **16**, 1133–1144 (2006).
63. Naoi, K. & Simon, P. New materials and new configurations for advanced electrochemical capacitors. *Electrochem. Soc. Interf.* **17**, 34–37 (2008).
64. Fischer, A. E., Pettigrew, K. A., Rolison, D. R., Stroud, R. M. & Long, J. W. Incorporation of homogeneous, nanoscale MnO₂ within ultraporos carbon structures via self-limiting electroless deposition: Implications for electrochemical capacitors. *Nano Lett.* **7**, 281–286 (2007).
65. Kazaryan, S. A., Razumov, S. N., Litvinenko, S. V., Kharisov, G. G. & Kogan, V. I. Mathematical model of heterogeneous electrochemical capacitors and calculation of their parameters. *J. Electrochem. Soc.* **153**, A1655–A1671 (2006).
66. Amatucci, G. G., Badway, F. & DuPasquier, A. in *Intercalation Compounds for Battery Materials* (ECS Proc. Vol. 99) 344–359 (Electrochemical Society, 2000).
67. Burke, A. R&D considerations for the performance and application of electrochemical capacitors. *Electrochim. Acta* **53**, 1083–1091 (2007).
68. Portet, C., Taberna, P. L., Simon, P. & Laberty-Robert, C. Modification of Al current collector surface by sol-gel deposit for carbon-carbon supercapacitor applications. *Electrochim. Acta* **49**, 905–912 (2004).
69. Talapatra, S. *et al.* Direct growth of aligned carbon nanotubes on bulk metals. *Nature Nanotech.* **1**, 112–116 (2006).
70. Taberna, L., Mitra, S., Poizot, P., Simon, P. & Tarascon, J. M. High rate capabilities Fe₃O₄-based Cu nano-architected electrodes for lithium-ion battery applications. *Nature Mater.* **5**, 567–573 (2006).
71. Jang, J. H., Machida, K., Kim, Y. & Naoi, K. Electrophoretic deposition (EPD) of hydrous ruthenium oxides with PTFE and their supercapacitor performances. *Electrochim. Acta.* **52**, 1733 (2006).
72. Cambaz, Z. G., Yushin, G., Osswald, S., Mochalin, V. & Gogotsi, Y. Noncatalytic synthesis of carbon nanotubes, graphene and graphite on SiC. *Carbon* **46**, 841–849 (2008).
73. Tsuda, T. & Hussey, C. L. Electrochemical applications of room-temperature ionic liquids. *Electrochem. Soc. Interf.* **16**, 42–49 (2007).
74. Balducci, A. *et al.* High temperature carbon–carbon supercapacitor using ionic liquid as electrolyte. *J. Power Sources* **165**, 922–927 (2007).
75. Balducci, A. *et al.* Cycling stability of a hybrid activated carbon/poly(3-methylthiophene) supercapacitor with *N*-butyl-*N*-methylpyrrolidinium bis(trifluoromethanesulfonyl)imide ionic liquid as electrolyte. *Electrochim. Acta* **50**, 2233–2237 (2005).
76. Balducci, A., Soavi, F. & Mastragostino, M. The use of ionic liquids as solvent-free green electrolytes for hybrid supercapacitors. *Appl. Phys. A* **82**, 627–632 (2006).
77. Endres, F., MacFarlane, D. & Abbott, A. (eds) *Electrodeposition from Ionic Liquids* (Wiley-VCH, 2008).
78. Faggioli, E. *et al.* Supercapacitors for the energy management of electric vehicles. *J. Power Sources* **84**, 261–269 (1999).
79. Chmiola, J. & Gogotsi, Y. Supercapacitors as advanced energy storage devices. *Nanotechnol. Law Bus.* **4**, 577–584 (2007).
80. Portet, C., Yushin, G. & Gogotsi, Y. Electrochemical performance of carbon onions, nanodiamonds, carbon black and multiwalled nanotubes in electrical double layer capacitors. *Carbon* **45**, 2511–2518 (2007).

Acknowledgements

We thank our students and collaborators, including J. Chmiola, C. Portet, R. Dash and G. Yushin (Drexel University), P. L. Taberna and C. Largeot (Université Paul Sabatier), and J. E. Fischer (University of Pennsylvania) for experimental help and discussions, H. Burnside (Drexel University) for editing the manuscript and S. Cassou (Toulouse) for help with illustrations. This work was partially funded through the Department of Energy, Office of Basic Energy Science, grant DE-FG01-05ER05-01, and through the Délégation Générale pour l'Armement.